



# PERCENTAGE MASS

You can find the percentage by mass of an element in a compound.

e.g.      % by mass of Fe in  $\text{Fe}_2\text{O}_3 = 100 \times \frac{2(56)}{160} = 70\%$

- 1    % by mass of O in  $\text{H}_2\text{O}$  .....
- 2    % by mass of O in  $\text{SO}_3$  .....
- 3    % by mass of Na in  $\text{Na}_2\text{CO}_3$  .....
- 4    % by mass of N in  $\text{Ca}(\text{NO}_3)_2$  .....
- 5    % by mass of O in  $\text{Ca}(\text{NO}_3)_2$  .....
- 6    % by mass of Cl in  $\text{FeCl}_3$  .....
- 7    % by mass of N in  $\text{NH}_4\text{NO}_3$  .....
- 8    % by mass of S in  $\text{Al}_2(\text{SO}_4)_3$  .....
- 9    % by mass of Li in lithium oxide .....
- 10   % by mass of Cr in chromium(III) oxide .....
- 11   % by mass of O in calcium hydroxide .....
- 12   % by mass of N in ammonium iodide .....
- 13   % by mass of O in  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$  .....

Area	Strength	To develop	Area	Strength	To develop	Area	Strength	To develop
Done with care and thoroughness			Can work out % by mass			Shows suitable working		
Can work out $M_r$			Can write formulae			Suitable and correct rounding		